

Communication

Subscriber access provided by ISTANBUL TEKNIK UNIV

Self-Assembly of a 3-D Triply Interlocked Chiral [2]Catenane

Aleema Westcott, Julie Fisher, Lindsay P. Harding, Pierre Rizkallah, and Michaele J. Hardie

J. Am. Chem. Soc., 2008, 130 (10), 2950-2951 • DOI: 10.1021/ja8002149

Downloaded from http://pubs.acs.org on February 8, 2009



More About This Article

Additional resources and features associated with this article are available within the HTML version:

- Supporting Information
- Links to the 2 articles that cite this article, as of the time of this article download
- Access to high resolution figures
- Links to articles and content related to this article
- Copyright permission to reproduce figures and/or text from this article

View the Full Text HTML





Published on Web 02/16/2008

Self-Assembly of a 3-D Triply Interlocked Chiral [2]Catenane

Aleema Westcott,[†] Julie Fisher,[†] Lindsay P. Harding,[‡] Pierre Rizkallah,[§] and Michaele J. Hardie^{*,†}

School of Chemistry, University of Leeds, Leeds LS2 9JT, U.K., Department of Chemical and Biological Sciences,

University of Huddersfield, Huddersfield HD1 3DH, U.K., and STFC Daresbury Laboratory, Daresbury,

Warrington WA4 4AD, U.K.

Received January 10, 2008; E-mail: m.j.hardie@leeds.ac.uk

Threaded molecules such as catenanes and rotaxanes have garnered much attention in recent years,¹ and concurrent with this has been the development of 3-D metallo-supramolecular assemblies that can be formed with a variety of prismatic and polyhedral structures.² Threaded metallo-supramolecular assemblies have been largely restricted to catenanes of 2-D metalla-cycles^{3,4} and include examples of doubly braided [2]catenanes.⁴ To the best of our knowledge, the only example of an interlocking 3-D metallosupramolecular assembly is the [2]catenane reported by Fujita and co-workers.⁵ In this assembly two asymmetric trigonal bipyramidal $[M_3(L1)(L2)]^{6+}$ cages interlock and the formation of the catenane is driven by $\pi - \pi$ stacking interactions between the two types of ligand. Despite this assembly being known for the best part of a decade, a second system displaying the same interlocking motif has yet to be reported. We report the second such system herein. This occurs with the cyclotriveratrylene-related molecular host (\pm) -2,7,12-trimethoxy-3,8,13-tris(4-[4'-methyl-2,2'-bipyridyl]benzyloxy)-10,15-dihydro-5*H*-tribenzo[*a*,*d*,*g*]cyclononene **1** in the complexes $[M_3(1)_2(NO_3)_3]_2^{6+}$ where M = Zn(II) or Co(II). Cyclotriveratrylene (CTV) has a relatively rigid bowl-shape, and a small number of single cage 3-D metallo-supramolecular assemblies with cyclotriveratrylene-based ligands have been previously reported.^{6,7}

Ligand 1 was synthesized in 64% yield from reaction of 4-bromomethyl-4'-methyl-2,2'-bipyridine with cyclotriguaiacylene in dry dimethylformamide in the presence of NaH. Two different cyclotriveratrylene-based ligands with appended 2,2'-bipyridine moieties have been previously reported.⁸ Ligand 1 self-assembles with $Zn(NO_3)_2$ or $Co(NO_3)_2$ in dimethylsulfoxide (DMSO) to form $[M_3(1)_2(NO_3)_3]_2 \cdot 6(NO_3)[2]$ catenanes, Scheme 1, which have been characterized by crystallography in the case of the Zn(II) complex and by mass spectrometry.

Very small octahedral crystals of $[Zn_3(1)_2(NO_3)_3]_2 \cdot 6(NO_3) \cdot$ 10(DMSO) 2 were grown in 63% yield by diffusing acetone vapors into a DMSO solution of ligand 1 and 2 equiv of $Zn(NO_3)_2$. The crystals were very weakly diffracting, however a single-crystal structure could be obtained by utilizing synchrotron radiation. Orange crystals with the same morphology were obtained using Co(NO₃)₂ in place of Zn(NO₃)₂, but even with synchrotron radiation, these did not show sufficient diffraction to establish a unit cell.

The asymmetric unit of complex 2 comprises a Zn(II) cation, a nitrate, and one-third each of two crystallographically distinct ligands. A second nitrate anion and solvent DMSO could not be located, and the given composition was established by elemental analysis. The Zn(II) has distorted octahedral geometry with one chelating nitrate anion and two chelating bpy moieties from the two distinct ligands. Each ligand coordinates to three equivalent Zn(II) centers creating a trigonal bipyramidal coordination cage with 15.96 Å between Zn(II) centers. Trigonal bipyramidal M₃L₂

Scheme 1. Self-Assembly of [2]Catenane Species



cages are known for a number of systems,^{5,9} including single cage examples with CTV-based ligands from Shinkai et al.6

A trigonal bipyramidal cage contains three windows and in complex 2 two such cages interlock through all three windows to form a [2]catenane; see Scheme 1 and Figure 1. The molecular bowls of the host ligands of different cages stack on top of one another in a slightly misaligned manner. There are no face-to-face $\pi - \pi$ stacking interactions evident between these molecular bowls, given a separation of 4.77 Å between the centers of the aromatic rings. Within the inner part of the [2]catenane two ligand 1 molecular bowls are facing one another creating an inner binding core. There is significant space within this inner core estimated at ~ 200 Å³.

The two cages of the catenane do not interlock in a centered fashion. This is due to two types of weak hydrogen bonding interaction that occur between the catenating cages. The closest interaction is between a CH of the bpy of one cage and the OMe groups of the internal ligands of the second cage at C- H···O distance 2.37 Å (corresponding C···O separation 3.17 Å). There is also C-H···O hydrogen bonding between the unbound O of the nitrate ligand and the pendant methyl group on the bpy moieties of one of the two types of ligand, at C-H···O distance 2.46 Å and C...O separation 3.37 Å. Both types of hydrogen bond occur at six positions in the core or around the periphery of the [2]catenane respectively, Figure 1.

A single $[Zn_3(1)_2(NO_3)_3]^{3+}$ cage contains both enantiomers of the ligand; however, within each [2]catenane all six tris-chelated

[†] University of Leeds.

[‡] University of Huddersfield. [§] STFC Laboratory.



Figure 1. From the crystal structure of $[Zn_3(1)_2(NO_3)_3]_2 \cdot 6(NO_3) \cdot 10(DMSO)$ 2, giving two views of the [2]catenane. Hydrogen atoms not involved in hydrogen bonding are excluded for clarity. The two catenating cages have different colors.

Zn(II) centers have the same hand and hence the [2]catenane is chiral. Overall, the complex crystallizes as a racemic mixture. There are extensive face-to-face $\pi - \pi$ stacking interactions between [2]catenane assemblies in the crystal lattice at an aryl ring centroid separation of 3.56 Å. Each $[Zn_3(1)_2(NO_3)_3]_2^{6+}$ forms $\pi - \pi$ interactions with six others to give an octahedral network of the [2]catenanes.

1-D ¹H NMR experiments were carried out on d_6 -DMSO solutions of 1 and Zn(NO₃)₂ in 2:3 and 1:3 proportions at various concentrations and temperatures but in all cases the spectra were too broad to assign. 2-D diffusion ordered (DOSY) NMR gave a molecular mass for the main component in solution consistent with a single [Zn₃(1)₂(NO₃)₃]³⁺ cage. Electrospray mass spectrometry (ES-MS) studies also show the presence of the single cage along with the [2]catenane through an overlapping doubly and singly charged peak at m/z 2416.55 which corresponds to {[Zn₃(1)₂- $(NO_3)_3]_2 \cdot (NO_3)_4\}^{2+}$ and $\{[Zn_3(1)_2(NO_3)_3] \cdot (NO_3)_2\}^+$ (calcd 2416.68). Additional peaks were present for species $\{[Zn_2(1)(NO_3)_2] \cdot (NO_3)\}^+$,

 $\{[Zn(1)(NO_3)]\}^+$, and $\{[Zn(1-C_{12}H_{11}N_2)(NO_3)]\}^+$. ES-MS studies of the Co(II)-system show the presence of the [2]catenane much more clearly with a triply charged peak at m/z 1573.04 corresponding to ${[Co_3(1)_2(NO_3)_3]_2 \cdot (NO_3)_3]^{3+}}$ (calcd 1575.33). Additional peaks were assigned to the 1:1 Co:1 species $\{[Co(1)(NO_3)]\}^+$ and $\{[Co(1-C_{12}H_{11}N_2)(NO_3)]\}^+$ and there was no indication of a single cage species. The observed loss of a methyl-bipyridine arm in species $\{[Zn(1-C_{12}H_{11}N_2)(NO_3)]\}^+$ and $\{[Co(1-C_{12}H_{11}N_2)(NO_3)]\}^+$ indicate that the assemblies are fragmenting in the mass spectrometer. It is also notable that, as might be expected for a [2]catenane, there are only [M₃L₂]₂, [M₃L₂], or smaller species observed.

The self-assembly of the $[M_3(1)_2(NO_3)_3]_2 \cdot 6(NO_3)$ [2]catenanes demonstrates that the triply interlocking [2]catenane motif is accessible for chemical systems other than Fujita's original example. Furthermore, it can be formed by symmetrical 3-D metallosupramolecular assemblies where the ligands do not show extensive $\pi - \pi$ stacking between them. Homometric or "self-stacking" motifs as seen in complex 2 are common with CTV and its derivatives^{7c,10} and may provide clues as to how this complicated topological motif occurs in this instance.

Acknowledgment. We thank Martin Huscroft for elemental analysis, and the CCLRC for access to microcrystal diffraction facilities. Funding from the EPSRC is gratefully acknowledged.

Supporting Information Available: Full experimental details; mass spectra; additional structure diagrams; crystal information in CIF format. This material is available free of charge via the Internet at http:// pubs.acs.org.

References

- Sauvage, J.-P., Dietrich-Buchecker, C., Eds. *Molecular Catenanes, Rotaxanes and Knots*; Wiley-VCH: Weinheim, 1999.
 For recent reviews, see: (a) Fujita, M.; Tominaga, M.; Hori, A.; Therrien, B. Acc. Chem. Res. 2005, 38, 369. (b) Seidel, S. R.; Stang, P. J. Acc. Chem. Res. 2002, 35, 972. (c) Swiegers, G. H.; Malefetse, T. J. Chem. P. 2009, 160, 2492. Rev. 2000, 100, 3483.
- (3) For examples, see: (a) Chas, M.; Pia, E.; Toba, R.; Peinador, C.; Quintela, J. M. *Inorg. Chem.* 2006, 45, 6117. (b) Hutin, M.; Schalley, C. A.; Bernardinelli, G.; Nitschke, J. R. *Chem. Eur. J.* 2006, 12, 4069. (c) Fuller, A.-M. L.; Leigh, D. A.; Lusby, P. J.; Slawin, A. M. Z.; Walker, D. B. J. Am. Chem. Soc. 2005, 127, 12612. (d) McArdle, C. P.; Van, S.; Jennings, MCC Hen. Soc. 2005, 127, 12612. (d) McArdle, C. P.; Van, S.; Jennings, MCC, P.; Van, S.; Jennings, P.; Jennings, MCC, P.; Van, S.; Jennings, P.; Jennings, MCC, P.; Van, S.; Jennings, P.; Jen M. C.; Puddephatt, R. J. J. Am. Chem. Soc. 2002, 124, 3959. (e) Hori, A.; Kumazawa, K.; Kusukawa, T.; Chand, D. K.; Fujita, M.; Sakamoto, S.; Yamaguchi, K. *Chem. Eur. J.* **2001**, *7*, 4142. (f) Padilla-Tosta, M. E.; Fox, O. D.; Drew, M. G. B.; Beer, P. D. Angew. Chem., Int. Ed. 2001,
- (4) McArdle, C. P.; Vittal, J. J.; Puddephatt, R. J. Angew. Chem., Int. Ed. 2000, 39, 3819.
- (5) Fujita, M.; Fujita, N.; Ogura, K.; Yamaguchi, K. Nature 1999, 400, 52. (6) Zhong, Z.; Ikeda, A.; Shinkai, S.; Sakamoto, S.; Yamaguchi, K. Org. Lett. 2001. 3. 1085.
- (a) Ronson, T. K.; Fisher, J.; Harding, L. P.; Hardie, M. J. Angew. Chem., Int. Ed. 2007, 46, 9086. (b) Sumby, C. J.; Hardie, M. J. Angew. Chem., Int. Ed. 2005, 44, 6395. (c) Sumby, C. J.; Fisher, J.; Prior, T. J.; Hardie,
- M. J. Chem. Eur. J. 2006, 12, 2945.
 (8) (a) Hardie, M. J.; Mills, R. M.; Sumby, C. J. Org. Biomol. Chem. 2004, 2, 2958. (b) Wytko, J. A.; Boudon, C.; J. Weiss, J.; Gross, M. Inorg. Chem. 1996, 35, 4469.
- (9) For examples, see: (a) Mukherjee, P. S.; Das, N.; Stang, P. J. J. Org. *Chem.* 2004, *69*, 3526. (b) Fujita, M.; Nagao, S.; Ogura, K. J. Am. Chem. *Soc.* 1995, *117*, 1649. (c) Sun, W.-Y.; Fan, J.; Okamura, T.; Xie, J.; Yu, K.-B.; Ueyama, N. *Chem. Eur. J.* 2001, *7*, 2557. (d) Bretonnière, Y.;
- Mazzanti, M.; Wietzke, R.; Pécaut, J. Chem. Commun. 2000, 1543.
 (10) For examples, see: (a) Wang, S.-Q.; Zeng, G.; Zheng, X.-F.; Zhao, K. Acta Crystallogr., Sect. E 2003, 59, 01862. (b) Scott, J. L.; MacFarlane, D. R.; Raston, C. L.; Teoh, C. M. Green Chem. 2000, 2, 123. (c) Steed, J. W.; Zhang, H.; Atwood, J. L. Supramol. Chem. 1996, 7, 37.

JA8002149